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INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(51) International Patent Classification ⁶ : C01B 3/40, 3/38, 3/48		A1	(11) International Publication Number: WO 99/48805 (43) International Publication Date: 30 September 1999 (30.09.99)
(21) International Application Number: PCT/GB99/00753 (22) International Filing Date: 23 March 1999 (23.03.99) (30) Priority Data: 9806199.7 24 March 1998 (24.03.98) GB		(81) Designated States: AE, AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, CA, CH, CN, CU, CZ, DE, DK, EE, ES, FI, GB, GD, GE, GH, GM, HR, HU, ID, IL, IN, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, UA, UG, US, UZ, VN, YU, ZA, ZW, ARIPO patent (GH, GM, KE, LS, MW, SD, SL, SZ, UG, ZW), Eurasian patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European patent (AT, BE, CH, CY, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE), OAPI patent (BF, BJ, CF, CG, CI, CM, GA, GN, GW, ML, MR, NE, SN, TD, TG). Published <i>With international search report.</i>	
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(54) Title: CATALYTIC GENERATION OF HYDROGEN (57) Abstract <p>A process for the catalytic generation of hydrogen by the self-sustaining combination of partial oxidation and steam reforming of a hydrocarbon comprises contacting a mixture of the hydrocarbon, an oxygen-containing gas and steam with a catalyst comprising rhodium dispersed on a refractory oxide support material which is a mixture of ceria and zirconia. The hydrocarbons are straight chain or branch chain hydrocarbons having 1 to 15 carbon atoms and include methane, propane, butane, hexane, heptane, normal-octane, iso-octane, naphthas, liquefied petroleum gas and reformulated gasoline petrol and diesel fuels. The hydrogen generation process can be started by feeding the hydrocarbon and air to initiate partial oxidation, before steam is added. The hydrogen generation process also may be operated in combination with a water-gas shift reaction for the reduction of carbon monoxide in the hydrogen generated.</p>			

CATALYTIC GENERATION OF HYDROGEN

This invention relates to the catalytic generation of hydrogen from hydrocarbons.

5 Hydrogen is currently used mainly in industry, in activities such as the manufacture of fertilisers, petroleum processing, methanol synthesis, annealing of metals and producing electronic materials. In the foreseeable future, the emergence of fuel cell technology will extend the use of hydrogen to domestic and vehicle applications.

10 A fuel cell works best when the anode is supplied with neat hydrogen. In the design of practical systems, however, other factors also need to be considered, including the availability, cost, supply, distribution, storage and release of clean hydrogen. When all these factors are taken into consideration, alternative methods of fuelling can show an overall advantage.

15 The issue of fuelling is very dependent on the type of application. For example, the design of fuel cell powered passenger vehicles requires a compact and responsive supply of hydrogen which must provide comparable driving performance to that of a combustion powered vehicle, as well as achieving higher efficiency and improved emission standards.

20 Although conventional and novel on-board hydrogen storage options are being developed, these do not seem likely to meet the target requirements for mass, size and cost, in time to be used for the first generation of fuel cell vehicles. Instead, the technology most likely to be implemented in the short term is the on-board generation of hydrogen from a liquid or liquefied fuel. On the other hand, the design of domestic systems for generating heat and fuel cell power is less constrained by the need for compactness and speed of response.

25 Furthermore, as the most widely available domestic fuel is natural gas, the efficient conversion of methane to hydrogen is seen as a key development target.

30 New fuel-processing technologies for generating hydrogen tend to be based either on steam-reforming or on partial oxidation. Each approach has its merits. Partial oxidation is a fast exothermic process, resulting in rapid start-up and short response times. Steam

reforming is endothermic and very efficient, producing hydrogen from both the fuel and the steam.

However, our work on system simulations leads us to predict that, in terms of efficiency, the ideal fuel-processor is likely to function by a combination of partial oxidation and steam reforming. In our previous work, (see EP 0217532; EP 0262947; WO 96/00186 and Platinum Metals Review, 1989, 33 (3) 118-127), we have shown that the two reactions can be carried out simultaneously in the same catalyst bed using a catalytic hydrogen generator which has become known as the HotSpot™ reactor. The process which takes place in the HotSpot reactor is a self-sustaining combination of exothermic partial oxidation and endothermic steam-reforming to produce a gas-stream containing mainly hydrogen, carbon dioxide and nitrogen, and is characterised by a low rate of carbon monoxide formation.

Possibly, the easiest fuel to process is methanol. The merits of methanol are well documented and include:

- (i) low propensity for soot formation;
- (ii) absence of contaminants (especially sulphur);
- (iii) possibility of production from renewable sources; and
- 20 (iv) availability of compatible components.

However, the disadvantages of methanol are equally familiar, notably:

- (i) relatively high toxicity;
- (ii) high affinity for water, resulting in corrosiveness;
- 25 (iii) absence of infrastructure for supplying vehicle fuelling stations; and
- (iv) unsuitability for domestic use.

Although there has been considerable progress in the development of on-board methanol steam reformers, because of the above-mentioned disadvantages of methanol, it 30 is by no means certain that methanol will be widely adopted by the manufacturers of fuel-cell vehicles. For domestic applications, natural gas is overwhelmingly the preferred choice.

The question of supply and distribution, in particular, has emerged as one of the key issues in the debate on the fuelling of fuel-cell systems, with a strong case being made for the use of the most widely available fuels. This caused us to further investigate the feasibility of generating hydrogen from hydrocarbon fuels by the self-sustaining reaction of air and steam as can be accomplished *inter alia* by our HotSpot reactor.

An object of the present invention, therefore, is to provide an improved process for generating hydrogen from hydrocarbons by a self-sustaining combination of partial oxidation and steam reforming.

10

In the course of our investigations we have identified a catalyst species which is extremely effective for oxidatively reforming (*ie* by partial oxidation alone or with the combination of steam reforming) hydrocarbons allowing hydrogen to be formed with high selectivity.

15

According to the present invention there is provided a process for the catalytic generation of hydrogen by the self-sustaining partial oxidation of a hydrocarbon comprising contacting a mixture of the hydrocarbon and an oxygen-containing gas and optionally steam with a catalyst comprising rhodium dispersed on a refractory oxide support material which comprises as cations cerium and zirconium.

Preferably, steam is introduced into the mixture of hydrocarbon and oxygen-containing gas after the self-sustaining partial oxidation of the hydrocarbon has commenced.

25

Further preferably, the hydrocarbon is a straight chain or branch chain hydrocarbon having 1 to 15 carbon atoms, conveniently 1 to 7 carbon atoms.

Suitably, the hydrocarbon is selected from methane, propane, butane, hexane, heptane, normal-octane, iso-octane, naphthas, liquified petroleum gas, reformulated petrol and diesel-type fuels.

Suitably, the oxygen-containing gas is air.

Preferably, rhodium comprises 0.1 weight% to 5 weight% of the total weight of the supported catalyst, conveniently 0.2 weight% to 2.5 weight%.

5

Preferably, the refractory oxide support material is a mixture of ceria and zirconia.

Preferably, the weight ratio of ceria to zirconia in the catalyst support material is from 0.5 to 99.5 to 99.5 to 0.5, conveniently from 5 to 95 to 95 to 5.

10

Preferably, the catalyst is pre-heated to a temperature at which self-sustaining partial oxidation of the hydrocarbon commences. The catalyst may be pre-heated by direct heating or catalytic heating.

15

Further preferably, the catalytic heating method comprises feeding to the catalyst an oxygen-containing gas and an initiating compound which is more easily oxidisable than the hydrocarbon to be partially oxidised, suitably, methanol, hydrogen or dimethyl ether.

20

Preferably, the mixture of hydrocarbon and oxygen-containing gas is fed to the catalyst when the catalyst has reached the temperature at which self-sustaining partial oxidation of the hydrocarbon will occur.

25

The present invention in a preferred form is a self-sustaining combination of partial oxidation and steam reforming of the hydrocarbon, the steam reforming being conducted by introducing steam into the mixture of hydrocarbon and oxygen-containing gas after the self-sustaining partial oxidation of the hydrocarbon has commenced.

30

The process of the present invention may be operated in combination with a catalysed water-shift reaction for the reduction of carbon monoxide in the hydrogen produced from the hydrocarbon.

The catalyst for the water-gas shift reaction preferably is a copper or iron based catalyst.

5 The water-gas shift reaction catalyst may be added to the rhodium based catalyst for the hydrogen generation reaction.

From another aspect, the present invention is hydrogen whenever prepared according to the processes defined above.

10 From yet another aspect, the present invention is a fuel cell system comprising a hydrogen production unit which operates according to the processes defined above.

15 We have found that a high proportion of zirconia induces a lower light-off temperature, but results in a higher temperature for self-sustaining operation and so causes more rapid deactivation of the catalyst whereas a high proportion of ceria lowers the temperature of self-sustaining operation and improves durability.

Advantages of the process of the present invention include the following:

- (i) it allows existing infrastructure straight chain hydrocarbon fuels to be used for the
20 generation of hydrogen;
- (ii) it allows for highly compact and responsive hydrogen generation;
- (iii) it is self-sustaining;
- (iv) it operates at comparatively low temperatures;
- (v) it produces mainly hydrogen and carbon dioxide without the need for one or more
25 separate or integral water shift reactors (and thus is a considerable improvement over existing fuel generation technologies);
- (vi) there is little or no carbon deposition on the catalyst;
- (vii) there is no evidence of sulphur poisoning of the catalyst; and
- (viii) it provides a catalyst system which shows high conversion efficiency and high
30 selectivity towards the formation of hydrogen.

The present invention is further described by way of the following illustrative Examples.

The following background information is given with respect to the illustrative
5 Examples.

(i) Micro-scale Reactor Temperature-programmed Experiments

A micro-reactor test rig was designed to measure the rate of hydrogen generation and
10 to provide a complete product analysis during:
(a) temperature-programmed reaction of each fuel with air and air/water; and
(b) self-sustaining oxidative reforming of each fuel.

For each fuel, optimum experimental conditions for partial oxidation and auto-
15 thermal reforming were calculated. The reactor was then furnace heated up to the initial
temperature to be studied; this was the minimum temperature required to maintain any liquid
feed(s) in the vapour phase. When the temperature had stabilised, the calculated flows were
introduced to the catalyst and the composition of the exit stream was analysed. The
reactants were always supplied in the gas phase, *i.e.* liquid feeds were pre-vaporised. The
20 furnace temperature was then increased incrementally until full conversion of the fuel was
achieved. When the conditions for maximum hydrogen generation were established, further
experiments were carried out to see if varying the air, fuel or water feeds would further
improve the hydrogen output.

25 (ii) Micro-scale Catalyst Stability Tests

After optimising the conditions for hydrogen generation for each catalyst/fuel
combination, a durability test was carried out for 6-8 hours under those conditions. The
reformate composition was recorded at hourly intervals. Decline in catalyst activity could
30 be monitored from any change in hydrogen output. After completion of the test, the catalyst
was inspected for any signs of carbon retention.

(iii) Micro-scale Reactor Self-sustaining Experiments and HotSpot Reactor Experiments

Light-off could be induced by raising the catalyst bed temperature either by direct heating (using the furnace) or by catalytic heating (by feeding hydrogen and air at ambient 5 temperature). When the catalyst bed reached the light-off temperature (known from the temperature-programmed experiments), the furnace was removed or the hydrogen/air feed was switched to fuel/air. When the calculated optimum temperature for auto-thermal operation was reached, water was introduced to the feed stream.

10 For scaled-up experiments using a HotSpot reactor, the basic method as outlined for micro-scale reactor self-sustaining experiments was employed.

(iv) Methane

15 This was used as a model for natural gas (See Examples 1 to 3).

(v) Straight chain Naphtha

Heptane was used as a model for straight run naphtha (see Example 4).

20

(vi) Reformulated Gasoline (RFG)

Reformulated gasoline (RFG) contains straight and branch chain hydrocarbons, aromatics, oxygenates and sulphur compounds but is composed mainly of straight chain 25 hydrocarbons (see Example 5).

(vii) Gasoline

Straight chain normal-octane and branch chain iso-octane were used as models for 30 gasoline (See Examples 6 and 7).

(viii) AVCAT

AVCAT is an aviation turbine fuel, which was used because of its compositional similarity to diesel fuel (See Example 8).

5

EXAMPLE 1
Temperature-programmed Methane Reforming

A batch of catalyst with a nominal composition of 1%Rh/CeO₂-ZrO₂ (based on the proportions of precursors) was prepared by impregnating 50g of 50:50 (by mass) ceria-zirconia support material with an aqueous Rh-salt solution. The required amount of impregnating solution (30cm³) was prepared by adding distilled water to 3.64g of aqueous rhodium(III) nitrate containing 0.5g rhodium.

15 The impregnating solution was added to the support material and mixed thoroughly. Excess water was removed from the resultant paste, which was then left for 2 hours to form a semi-rigid cake. After breaking-up the cake, the lumps were dried at 120°C for 8 hours, before being calcined in static air at 500°C for 2 hours. Finally, the catalyst was crushed, pelletised (using a pressure of 8500kg cm⁻² for 15 min) and sieved to collect granules in the 20 range 0.3-0.8mm diameter. No special activation was required prior to testing.

25 A small bed (0.2g) of the granulated catalyst prepared as above was loaded into a tubular quartz reactor, which was positioned at the centre of a furnace. A mixture of methane (9.5 standard cm³min⁻¹), air (25 standard cm³min⁻¹) and steam (31 standard cm³min⁻¹) was passed through the catalyst bed, while the temperature inside the furnace was raised at 2°C min⁻¹ from 105°C to 800°C. The rate of hydrogen production peaked when the catalyst bed temperature reached 555°C. At this temperature, 98.5% of the methane was converted to a reformate containing 21% H₂, 6% CO₂ 0.9% CO and 65% N₂ (plus water and unreacted methane).

30

EXAMPLE 2**Self-sustaining Methane Reforming on Micro-reactor Scale**

A bed (1.0g) of the catalyst prepared in Example 1 was loaded into a quartz reactor.

5 Unlike Example 1, however, no external heat was supplied by the furnace in this Example 2. Instead, the catalyst bed was initially heated by feeding hydrogen (200 standard $\text{cm}^3\text{min}^{-1}$) and air (174 standard $\text{cm}^3\text{min}^{-1}$). When the bed temperature reached 600°C, the gas-feed was switched to methane (40 standard $\text{cm}^3\text{min}^{-1}$) and air (174 standard $\text{cm}^3\text{min}^{-1}$). Under these partial oxidation conditions, the bed temperature stabilised at 625°C, and the methane

10 conversion reached 45%. The composition of the reformat was 10.5% H₂, 9% CO₂, 1.8% CO, 12% CH₄ and 60% N₂ (plus water).

When the gas-feed was changed to methane (24 standard $\text{cm}^3\text{min}^{-1}$), air (130 standard $\text{cm}^3\text{min}^{-1}$) and steam (124.5 standard $\text{cm}^3\text{min}^{-1}$), the bed temperature dropped 15 to 605°C, but the H₂ concentration in the reformat rose to 12.5%. Under these conditions, the catalyst was functioning by a self-sustaining combination of partial oxidation and steam reforming of methane. Furthermore, by reducing the heat loss from the reactor, the amount of steam reforming could be increased. This resulted in 97% conversion of methane, and a reformat containing 24% H₂, 11.5% CO₂, 0.8% CO and 49% N₂ (plus water and unreacted methane). There was no sign of de-activation during 7 hours of testing.

EXAMPLE 3**HotSpot Reforming of Methane**

25 A radial bed (80g) of the catalyst prepared in Example 1 was tested using a suitably modified HotSpot reactor. The HotSpot reactor allows multiple injection of fuel, water and air in a catalyst bed which catalyses partial oxidation and steam reforming of the fuel. Based on the results of Example 2, the HotSpot reactor was modified to include thermal insulation (to prevent heat losses from the reactor) and high-temperature connectors and fittings (to tolerate the high temperatures compared to methanol reforming).

The temperature of the radial catalyst bed was raised by feeding hydrogen (0.85 standard litre min⁻¹) and air (2.6 standard litre min⁻¹). When the bed temperature reached 600°C, the gas-feed was switched to methane (1.14 standard litre min⁻¹) and air (4.67 standard litre min⁻¹). The catalyst temperature remained at 600°C during the partial oxidation of methane. Dry analysis of the reformat, by gas chromatography and non-dispersive IR, showed 15% H₂, 3% CO₂, 2% CO, 66% N₂ and 14% CH₄.

When the water was added to the gas feed (by vaporising liquid at a rate of 4.6cm³ min⁻¹), the HotSpot reactor began to function by a self-sustaining combination of partial oxidation and steam reforming. This had the effect of lowering the bed temperature (to 540°C), but raising methane conversion to 97% and the rate of hydrogen production (to 135 litre h⁻¹). Dry analysis of the reformat now showed 32% H₂, 13% CO₂, 1.3% CO, 53% N₂ and 0.8% CH₄. When the feed rates were changed (to methane : 6.49 standard litre min⁻¹; air : 18.74 standard litre min⁻¹; steam : 6.31 standard litre min⁻¹), the hydrogen output rose to 585 litre h⁻¹, and remained stable during 7 hours of testing.

EXAMPLE 4

Self-sustaining Heptane Reforming on Micro-reactor Scale

A batch of catalyst was prepared by the method described in Example 1, except that the support material was 80:20 (by mass) ceria-zirconia.

A bed (0.2g) of catalyst prepared as above, was loaded into a quartz reactor. The catalyst bed temperature was raised to 200°C by furnace heating. The furnace was then switched off, before heptane vapour (3.8cm³min⁻¹) and air (64.5cm³min⁻¹) were passed through the catalyst bed. When the catalyst reached 575°C, steam was added (124.4cm³min⁻¹) and the air feed-rate was decreased (to 24.1cm³min⁻¹). The catalyst temperature stabilised at 625°C. The hydrogen concentration of the dried reformat was 22%.

EXAMPLE 5**Self-sustaining Reformulated Gasoline on Micro-reactor Scale**

A fresh bed (0.2g) of the catalyst prepared in Example 4 was loaded into a quartz reactor. The catalyst bed temperature was raised to 200°C by furnace heating. The furnace was then switched off, before reformulated gasoline vapour (produced by vaporising the liquid at a rate of 1.5cm³ hour⁻¹) and air (62.8cm³ min⁻¹) were passed through the catalyst bed. When the catalyst reached 600°C, steam was added (62.5cm³ min⁻¹), and the bed temperature stabilised at 590°C. The hydrogen concentration of the dried reformat was 10 28.5%.

EXAMPLE 6**Temperature-programmed Reforming of Normal-Octane**

15 A bed (0.2g) of the catalyst prepared in Example 1 was loaded into a quartz reactor, which was positioned at the centre of a furnace. A mixture of normal-octane vapour (produced by vaporising the liquid at a rate of 4cm³ hour⁻¹), air (175cm³ min⁻¹) and steam (produced by vaporising water at a rate of 4cm³ hour⁻¹) was passed through the catalyst bed, while the temperature inside the furnace was raised at 2°min⁻¹ from 400°C to 650°C. 20 The rate of hydrogen production reached a plateau when the catalyst bed temperature reached 550°C. At this temperature, all the normal-octane was converted to a reformat, which when dried contained 37%H₂, 12%CO₂, 7%CO and nitrogen. There was no sign of de-activation, when the catalyst bed temperature was maintained at 550°C for 4 hours.

EXAMPLE 7**Temperature-programmed Reforming of Iso-Octane**

The experimental procedure described in Example 6 above was followed exactly, except that the normal-octane was replaced by iso-octane. Again, the rate of hydrogen 30 production reached a plateau when the catalyst bed temperature reached 550°C. At this temperature, all the iso-octane was converted to a reformat, which when dried contained

33%H₂, 15%CO₂, 5%CO and nitrogen. There was no sign of de-activation, when the catalyst bed temperature was maintained at 550°C for 4 hours.

EXAMPLE 8

5

Temperature-programmed Reforming of AVCAT

A bed (0.2g) of the catalyst prepared in Example 1 was loaded into a quartz reactor, which was positioned at the centre of a furnace. A mixture of AVCAT fuel (produced by vaporising the liquid at a rate of 4cm³ hour⁻¹), air (300cm³ min⁻¹) and steam (produced by 10 vaporising water at a rate of 4cm³ hour⁻¹) was passed through the catalyst bed, while the temperature inside the furnace was raised at 2°min⁻¹ from 400°C to 650°C. The rate of hydrogen production reached a plateau when the catalyst bed temperature reached 600°C. At this temperature, most of the AVCAT fuel was converted to a reformate, which when dried contained 28%H₂, 14%CO₂, 3%CO and nitrogen. When held at this temperature, 15 there was some de-activation during the first hour, but then the output stabilised at 24%H₂, 14%CO₂, 3%CO and nitrogen.

CLAIMS

1. A process for the catalytic generation of hydrogen by the self-sustaining partial oxidation of a hydrocarbon comprising contacting a mixture of the hydrocarbon and an oxygen-containing gas and optionally steam with a catalyst comprising rhodium dispersed on a refractory oxide support material which comprises as cations cerium and zirconium.
5
2. A process according to claim 1 wherein steam is introduced into the mixture of hydrocarbon and oxygen-containing gas after the self-sustaining partial oxidation of the hydrocarbon has commenced.
10
3. A process according to claim 1 or 2 wherein the hydrocarbon is a straight chain hydrocarbon or a branch chain hydrocarbon.
15
4. A process according to claim 3 wherein the hydrocarbon contains 1 to 15 carbon atoms.
15
5. A process according to claim 4 wherein the hydrocarbon contains 1 to 7 carbon atoms.
20
6. A process according to any one of the preceding claims wherein the hydrocarbon is selected from methane, propane, butane, hexane, heptane, normal-octane, iso-octane, naphthas, liquified petroleum gas, reformulated petrol and diesel-type fuels.
25
7. A process according to any one of the preceding claims wherein the oxygen-containing gas is air.
30
8. A process according to any one of the preceding claims wherein rhodium comprises 0.1 weight *per cent* to 5 weight *per cent* of the total weight of the supported catalyst.

- 9 A process according to claim 8 wherein rhodium comprises 0.2 weight *per cent* to 2.5 weight *per cent* of the total weight of the supported catalyst.
10. A process according to any one of the preceding claims wherein the refractory oxide support material is a mixture of ceria and zirconia.
- 5
11. A process according to claim 10 wherein the weight ratio of ceria to zirconia in the catalyst support material is from 0.5 to 99.5 to 99.5 to 0.5
- 10 12. A process according to claim 11 wherein the weight ratio of ceria to zirconia in the catalyst support material is from 5 to 95 to 95 to 5.
- 15 13. A process according to any one of the preceding claims wherein the catalyst is pre-heated to a temperature at which self-sustaining partial oxidation of the hydrocarbon commences.
14. A process according to claim 13 wherein the catalyst is pre-heated by direct heating to a temperature at which self-sustaining partial oxidation of the hydrocarbon commences.
- 20 15. A process according to claim 13 wherein the catalyst is pre-heated by catalytic heating to a temperature at which self-sustaining partial oxidation of the hydrocarbon commences.
- 25 16. A process according to claim 15 wherein the catalyst is pre-heated by feeding to the catalyst an oxygen-containing gas and an initiating compound which is more easily oxidisable than the hydrocarbon which is to be partially oxidised.
17. A process according to claim 16 wherein the initiating compound is selected from methanol, hydrogen and dimethyl ether.

18. A process according to any one of the preceding claims wherein the mixture of hydrocarbon and oxygen-containing gas is fed to the catalyst when the catalyst has reached the temperature at which self-sustaining partial oxidation of the hydrocarbon will occur.
- 5 19. A process as claimed in any one of the preceding claims operated in combination with a catalysed water-gas shift reaction for the reduction of carbon monoxide in the hydrogen produced from the hydrocarbon.
- 10 20. A process as claimed in claim 19 wherein the catalyst for the water-gas shift reaction is a copper or iron based catalyst.
21. A process according to claim 19 or 20 wherein the water-gas shift reaction catalyst is added to the rhodium based catalyst for the hydrogen generation reaction.
- 15 22. Hydrogen whenever prepared by a process according to any one of the preceding claims.
23. A fuel cell system comprising a hydrogen production unit which operates according to the process of any one of claims 1 to 21.

INTERNATIONAL SEARCH REPORT

Intell. Pat Application No
PCT/GB 99/00753

A. CLASSIFICATION OF SUBJECT MATTER
IPC 6 C01B3/40 C01B3/38 C01B3/48

According to International Patent Classification (IPC) or to both national classification and IPC

B. FIELDS SEARCHED

Minimum documentation searched (classification system followed by classification symbols)
IPC 6 C01B

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

C. DOCUMENTS CONSIDERED TO BE RELEVANT

Category *	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
X	EP 0 548 679 A (IDEMITSU KOSAN CO) 30 June 1993 see page 3, line 12 - line 44 see page 4, line 20 - line 28 see page 5, line 33 - page 6, line 5 ---	1,3-12, 22,23
X	EP 0 640 561 A (SHELL INT RESEARCH) 1 March 1995 see claims; example 4 ---	1,3-12, 22,23
A	EP 0 495 534 A (IGARASHI AKIRA C O KOHGAKUIN U) 22 July 1992 see the whole document ---	1,3-12, 22,23
A	US 4 501 823 A (MASUDA KOJI) 26 February 1985 see claims ---	1,6, 10-12, 22,23
		-/-

Further documents are listed in the continuation of box C.

Patent family members are listed in annex.

* Special categories of cited documents :

- *A* document defining the general state of the art which is not considered to be of particular relevance
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- *&* document member of the same patent family

Date of the actual completion of the international search

9 June 1999

Date of mailing of the international search report

16/06/1999

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Internat'l Application No
PCT/GB 99/00753

C.(Continuation) DOCUMENTS CONSIDERED TO BE RELEVANT

Category	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	WO 96 16737 A (SHELL CANADA LTD ;SHELL INT RESEARCH (NL)) 6 June 1996 see examples 5,7	1,13-18
A	EP 0 199 878 A (WESTINGHOUSE ELECTRIC CORP) 5 November 1986 see page 6, line 16 – page 7, line 18	19-21

INTERNATIONAL SEARCH REPORT

Information on patent family members

Internatinal Application No

PCT/GB 99/00753

Patent document cited in search report	Publication date	Patent family member(s)			Publication date
EP 0548679	A 30-06-1993	AU 661742	B	03-08-1995	
		AU 3002092	A	01-07-1993	
		CA 2085315	A	21-06-1993	
		JP 5221602	A	31-08-1993	
		NZ 245394	A	28-03-1995	
EP 0640561	A 01-03-1995	AT 173233	T	15-11-1998	
		AU 671232	B	15-08-1996	
		AU 7032894	A	09-03-1995	
		BR 9403321	A	11-04-1995	
		CA 2129798	A	25-02-1995	
		CN 1106357	A	09-08-1995	
		DE 69414494	D	17-12-1998	
		DE 69414494	T	20-05-1999	
		JP 7309601	A	28-11-1995	
		NO 943107	A	27-02-1995	
		NZ 264173	A	26-09-1995	
		SG 47038	A	20-03-1998	
		ZA 9406353	A	22-03-1995	
EP 0495534	A 22-07-1992	JP 2043950	A	14-02-1990	
		JP 6004135	B	19-01-1994	
		JP 2002879	A	08-01-1990	
		JP 2079807	C	09-08-1996	
		JP 7090169	B	04-10-1995	
		CA 1333278	A	29-11-1994	
		DE 68916285	D	28-07-1994	
		EP 0333037	A	20-09-1995	
		US 5130114	A	14-07-1992	
US 4501823	A 26-02-1985	JP 59199042	A	12-11-1984	
		DE 3415634	A	31-10-1984	
WO 9616737	A 06-06-1996	AU 689996	B	09-04-1998	
		AU 4259296	A	19-06-1996	
		CA 2205178	A	06-06-1996	
		EP 0794834	A	17-09-1997	
		NO 972402	A	27-05-1997	
		ZA 9510014	A	04-06-1996	
EP 0199878	A 05-11-1986	DE 3584245	A	31-10-1991	
		JP 61247601	A	04-11-1986	